

# Covalent Bonding Chapter 8 Answer Key

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covalent bonds. The majority of covalent bonds form between nonmetallic elements. 8. Describe how the octet rule applies to covalent bonds. Atoms share valence electrons; the shared electrons complete the octet of each atom. 9. Illustrate the formation of single, double, and triple covalent bonds using Lewis structures.

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Single Covalent Bonds When only one pair of electrons is shared, such as in a hydrogen molecule, it is a single

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covalent bond. The shared electron pair is often referred to as the bonding pair. For a hydrogen molecule, shown in Figure 8.4, each covalently bonded atom equally attracts the pair of shared electrons.

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Single Covalent Bonds When only one pair of electrons is shared, such as in a hydrogen molecule, it is a single covalent bond. The shared electron pair is often referred to as the bonding pair. For a hydrogen molecule, shown in Figure 8.4, each covalently bonded atom equally attracts the pair of shared electrons.

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Chapter 8: Covalent Bonding - Mr. Miller's Classes

Chapter 8 Covalent Bonding and Molecular Structure 8-3 There are two types of repulsive forces between the two atoms. First, the nuclei repel because they are both positively charged. Second, the electrons repel because they are both negatively charged. The attractive forces between the two atoms result from the

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Chapter 8: Covalent Bonding and Molecular Structure

Chapter 8- Covalent/Ionic Bonds.

Monatomic Ions. Ionic Bonding.

Covalent Bonding. Molecule. consists of single atoms. Attraction of two oppositely charged ions. a bond that results from the sharing of valance electrons. neutral group of atoms

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joined together by covalent bonds.

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processes (or mechanisms) by which reactions occur to create molecular structures. It then describes some of the many ways these reactions make new compounds, examined by functional groups and corresponding common reaction mechanisms.

Throughout, this book includes biochemical and pharmaceutical examples with varying degrees of difficulty, with worked answers and without, as well as advanced topics in later chapters for optional coverage. Incorporates valuable and engaging applications of the content to biological and industrial uses Includes a wealth of useful figures and problems to support reader comprehension and study Provides a high quality chapter on stereochemistry as well as advanced topics such as synthetic polymers and spectroscopy for class

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